Average Atomic Mass

- The atomic mass values on the periodic table are a weighted average of all the isotopes of a certain element. The weighting includes the % abundance and mass of the atom.

Ex. Silver  
Ag-107  51.83 % abundance  
Ag-109  48.17 % abundance

For each isotope, divide the % abundance by 100 and then multiply by the mass number. Then add the result of each isotope together.

\[
AAM = (107 \times 0.5183) + (109 \times 0.4817) \\
= 107.97
\]