14. ANS: D PTS: 1 DIF: L2 REF: p. 222
OBJ: 8.1.1 Identify the information a molecular formula provides.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

15. ANS: A PTS: 1 DIF: L1 REF: p. 227
OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

16. ANS: A PTS: 1 DIF: L1 REF: p. 223
OBJ: 8.1.1 Identify the information a molecular formula provides.
STA: C.3.1 | C.3.2 | C.3.3 BLM: knowledge

17. ANS: C PTS: 1 DIF: L2 REF: p. 226
OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

18. ANS: B PTS: 1 DIF: L2 REF: p. 226
OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

19. ANS: D PTS: 1 DIF: L2 REF: p. 230
OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

20. ANS: B PTS: 1 DIF: L2 REF: p. 227
OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: application

OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

22. ANS: A PTS: 1 DIF: L2 REF: p. 226
OBJ: 8.1.2 Describe the representative units that define molecular compounds and ionic compounds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

23. ANS: D PTS: 1 DIF: L2 REF: p. 238 | p. 244
OBJ: 8.3.1 Describe the relationship between atomic and molecular orbitals.
STA: C.3.1 | C.3.2 | C.3.3 | C.3.4 BLM: application

OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: application

25. ANS: B PTS: 1 DIF: L2 REF: p. 230
OBJ: 8.2.1 Explain the result of electron sharing in covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: application

26. ANS: B PTS: 1 DIF: L1 REF: p. 224
OBJ: 8.1.2 Describe the representative units that define molecular compounds and ionic compounds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: knowledge

27. ANS: D PTS: 1 DIF: L2 REF: p. 232
OBJ: 8.2.2 Describe how coordinate covalent bonds are different from other covalent bonds.
STA: C.3.1 | C.3.2 | C.3.3 BLM: comprehension

28. ANS: A PTS: 1 DIF: L2 REF: p. 233
OBJ: 8.2.4 Explain how the strength of a covalent bond is related to its bond dissociation energy.
STA: C.3.1 | C.3.2 | C.3.3 BLM: application

OBJ: 8.2.4 Explain how the strength of a covalent bond is related to its bond dissociation energy.
STA: C.3.1 | C.3.2 | C.3.3 BLM: application
Network solids are substances in which all of the atoms are covalently bonded to each other. Melting these substances requires breaking covalent bonds throughout the solid. Two examples are diamond and silicon carbide.

PTS: 1  DIF: L2  REF: p. 252
OBJ: 8.4.3 Explain why the properties of covalent compounds are so diverse.
STA: C.3.1 | C.3.2 | C.3.3  BLM: comprehension