Crystalline structure:

Crystal lattice:

In the case of NaCl, the unit cell is said to be face-centered cubic. Polymer molecules are very large so it might seem that they could not pack together regularly and form a crystal. It now is known that regular polymers may form lamellar (plate-like) crystals with a thickness of 10 to 20 nm in which the parallel chains (shown in different colors in the simulated structure below) are perpendicular to the face of the crystals.

Lattices can either be primitive (only one lattice point per unit cell) or non-primitive (more than one lattice point per unit cell).

Crystalline solids contain highly ordered arrays of molecules and atoms held together by non-covalent interactions. The crystals of a given form may vary in size, the relative development of the given faces (forms) present, that is, they may have different crystal habits.

Crystal Habit:

The habit describes the overall shape of the crystal. For example acicular (needle like), prismatic, pyramidal, tabular, equant, columnar and lamellar types.