Hydrogen atom

Hydrogen molecule

e.g. In molecular bond in hydrogen

- Share want to achieve the structure of noble gas
- Can be similar to quartet structures
- Use molecular model by sharing electrons - a pair of electrons make one bond

Bonding

[Diagram of electron configurations]

- Chlorine ion
- Chlorine
- FL [Cl, 18]
- FL [2, 18, 8]

Chlorine atom chlorine from

- Chlorine ion makes a molecule by transferring an electron and making a covalent bond
- The nucleus loses electrons, non-molecular can detach
- There is nothing by non-localising strong forces of attraction between the oppositely charged ions
- Non-localised chlorine, non-molecular can detach
- Chlorine atom is actually formed when molecules react with non-molecules

Ionic Bonding

- Form by covalent bonds of atoms
- Bond formed when atoms are negatively charged and have different charges
- Long distances are necessary to separate the number enough to hold together and have different charges
- Can be found in compounds of chlorine, as in salt, while another such

As you descend down a group, you add another small chlorine atom number tells you how many electrons these are in the outer shell

This group number tells us what a main metal element have, the outer shell full or not. The number tells us the number of outer metal, each called represents a number of electrons