Enthalpy Changes in Solutions

- $\Delta H^\circ_{\text{sol}}$ is the change in enthalpy when 1 mole of a substance is dissolved in a large excess of a pure solvent

- $\Delta H_{\text{sol}} = \Delta H_{\text{lattice enthalpy}} + \Delta H_{\text{hydration}}$
  - Be careful: two ions for $\Delta H_{\text{hydration}}$

- Hydration: when water

- Solvation: anything else