**Chemical Bond** - A force of attraction that holds two atoms together.

**Ionic Bond** - A force of attraction between positively charged ions.

**Covalent Bond** - Electrons are shared between atoms.

**Molecule** - Neutral group of atoms held together by covalent bonds.

**Metallic Bond** - A force of attraction between a positively charged metal ion and the electrons in a metal.

**Physical Science Chapter 4 pages 92-99**

**Pure Substance** - A substance where here is only one type of particle.

**Metal** - Elements that are shiny and are good conductors.

**Nonmetal** - Elements that are not shiny and are poor conductors.

**Metalloid** - Elements that are semiconductors. Properties of metals, nonmetals, and metalloids.

**Compound** - Pure substance that is composed of two or more elements that are chemically combined.

**Mixture** - Combo of two or more substances that are not chemically combined.

**Solution** - A solution that looks to be a single substance that is composed of multiple substances.

**Homogeneous Mixture** - Is made of different substances that remain physically separate.

**Heterogeneous Mixture** - Has the same uniform appearance and composition throughout.

Many homogeneous mixtures are commonly referred to as solutions.

- Substance that is dissolved.

**Solute** - Substance in which the solute is dissolved.

**Concentration** - Amount of solute dissolved into a solvent.

**Solubility** - The property of a solid, liquid, or gaseous chemical substance called solute to dissolve in a solid, liquid, or gaseous solvent to form a solution of the solute in the solvent.

**Suspension** - Heterogeneous mixture containing solid particles that are sufficiently large for sedimentation. Usually they must be larger than one micrometer.

**Colloid** - A solution that has particles ranging between 1 and 1000 nanometers in diameter, yet still are able to remain evenly distributed throughout the solution.