Molecular geometry: based off the electronic geometry

Octahedral (8 faces)

Polarity: NaCl is ionic so you see a complete transfer of electrons. Hal and Cl₂ are covalent so you see a sharing of electrons. HCl is polar covalent and Cl₂ is non-polar covalent.

Molecular shapes based on tetrahedral electron-group geometry of CH₄, NH₃, and H₂O

**FIGURE 10-11**

Polarity: NaCl is ionic so you see a complete transfer of electrons. Hal and Cl₂ are covalent so you see a sharing of electrons. HCl is polar covalent and Cl₂ is non-polar covalent.