**Chemical Bonding**

**Ionic bond** – bond between inorganic ions (e.g. Na⁺ and Cl⁻)

**Covalent bond** – bond between atoms in molecules (e.g. CH₄ and H₂O), may be single, double, or triple bond.

- may be equal distribution of charge (CH₄)… non-polar… hydrophobic

- may be unequal distribution of charge (H₂O) (a.k.a. dipole)… polar… hydrophilic

**Disulfide bond** – bond between Sulphur atoms Important in protein folding for proper configuration and function. Cys=Cys molecules determine shape of molecule.

**Hydrogen bond** – weak attraction between oppositely charged atoms in a dipole, or between dipoles (particularly in the presence of -OH, -C=O, >N-H groups)