

This paper is suspended in a cylinder-shaped, hollow glass container. There is a shallow layer of liquid at the bottom. It must be below the pencil line on the paper. The container must be covered to ensure, that the atmosphere in the container is saturated with the vapour of the liquid and that it does not escape. Saturating the atmosphere in the beaker with vapour stops the solvent from evaporating as it rises up the paper.



You can then identify the different components of each liquid. This can be used to determine if a product is pure or not.

Chromatography results:

Rf value = Distance travelled / Solvent front

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