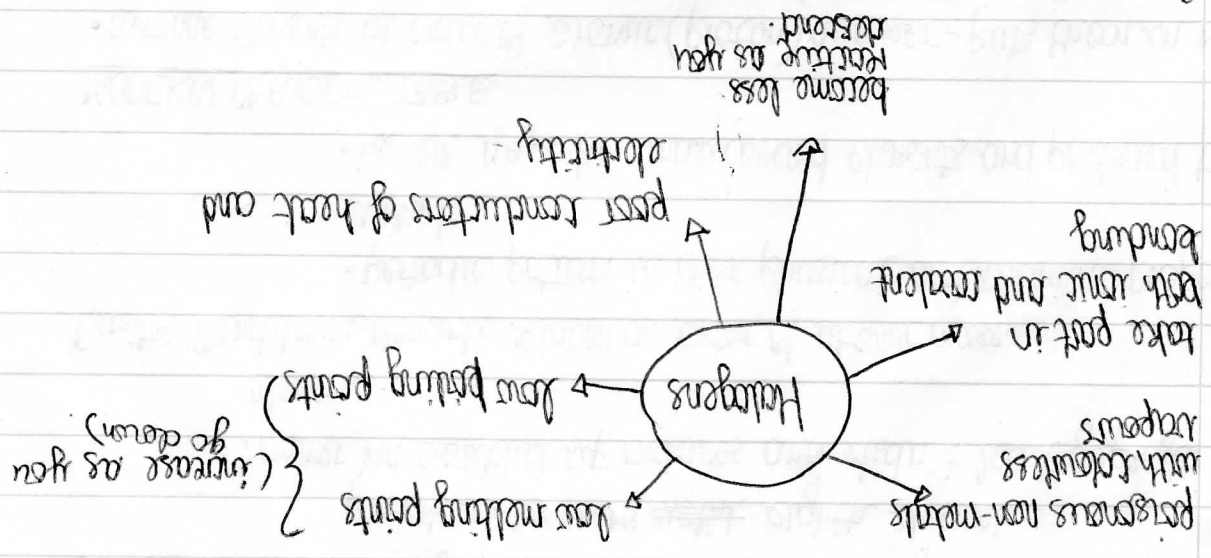




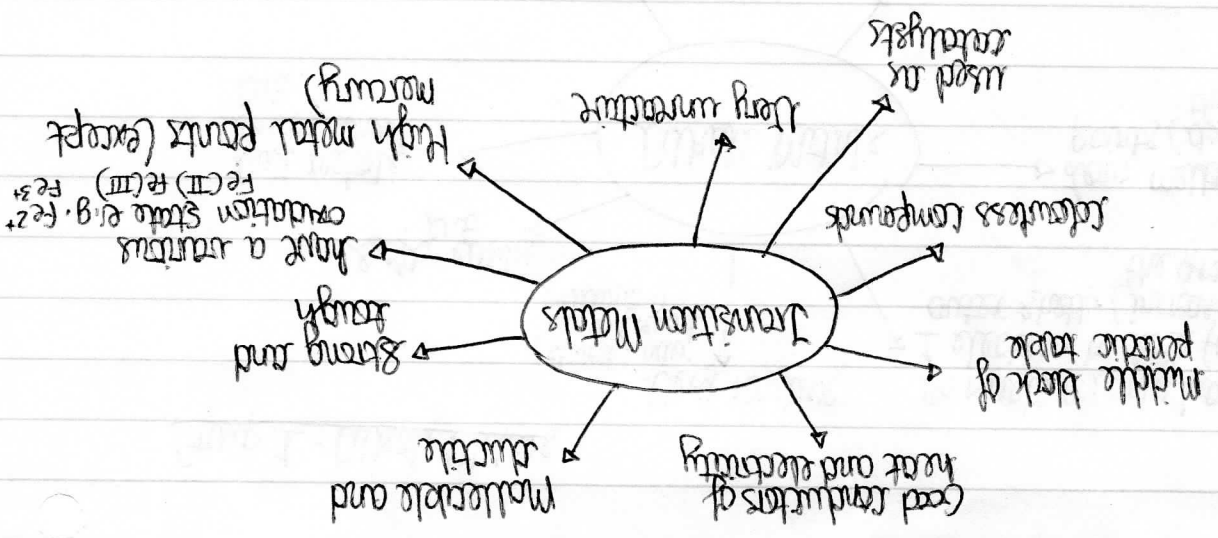
- \* React with metals to produce ionic salts (metal halides)
- \* React with non-metallic elements to form molecular compounds.
- \* more reactive halogen will displace less reactive halogen from an aqueous solution

ANALYSIS TEST

Add Silver Nitrate ( $\text{AgNO}_3$ ) solution + Nitric Acid ( $\text{HNO}_3$ ) forms precipitate:

|                        |                      |
|------------------------|----------------------|
| Chloride $\text{Cl}^-$ | - white precipitate  |
| Bromide $\text{Br}^-$  | - yellow precipitate |
| Iodide $\text{I}^-$    | - yellow precipitate |

Transition Elements (Metals)



\* 3rd energy level can hold up to 18 electrons once 2 electrons have occupied 4th energy level.