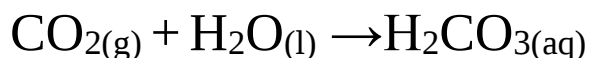


pH OF ACID RAIN

Normal rainwater is slightly acidic, with a pH of around 5.6, due to the dissolution of carbon dioxide (CO₂) in water, forming weak carbonic acid (H₂CO₃):



Acid rain has a pH below 5.6, typically ranging from 4.2 to 4.4, but it can be even lower in heavily polluted areas.

SOURCES OF ACID RAIN

Natural Sources

- Volcanic eruptions release sulfur dioxide and other gases.
- Decomposition of organic matter produces nitrogen oxides.
- Lightning generates nitrogen oxides.

Anthropogenic (Human-Made) Sources

- Burning of fossil fuels (coal, oil, and natural gas) in power plants, industries, and vehicles.
- Industrial processes, such as smelting of metal ores.
- Agricultural activities, including the use of fertilizers.