Atomic × A Z + Atomic Number

$$A = n + p$$

$$Z = p = e$$

0

Electron configuration

1st shell -> ze

2nd shell -> maxotesale.co.uk

3nd shell fromax. of 4

Preview page 2 of 4

Reactivity Group increases as Group you go down 1 because the electrons are further away the nucleus the nucleus so they are given up more easily

Reactivity inocease as you go up because of the same thing as in group 1