LE CHATELIER'S PRINCIPLE AND CHANGE IN TEMPERATURE

A change in temperature will cause the position of the equilibrium to change.
Temperature affects a reaction.
For example, **when temperature increases, the flow of heat increases into the system. Conversely, a decrease in temperature is the result of heat flowing out of the system.**

Enthalpy of Reaction describes the heat flow for a reaction.

If $\Delta H_{rxn} > 0$ then the reaction is endothermic, as the reaction draws heat from its surroundings.

If $\Delta H_{rxn} < 0$, the reaction is exothermic as heat is released into the surroundings as the reaction is occurring.

According to the Chatelier's Principle, the system will react to remove added heat to restore equilibrium by converting the products back to reactants.
However if heat is decreased, then the system will react in the opposite direction, causing the reaction to release heat into the system to restore the temperature.