

Corrosion

The slow process of decay and destruction of metals due to the action of air, moisture or acids is called corrosion.

For example:

Iron combines with oxygen present in the air, in the presence of water, to form a red-brown flaky substance called **rust**. This process is commonly called the **rusting of iron**.

The chemical formula of rust is $\text{Fe}_2\text{O}_3 \cdot x \text{H}_2\text{O}$.

Prevention of Corrosion

- Corrosion damages buildings, bridges, ships, automobiles and other articles made of iron. Hence, prevention of corrosion is necessary. This will not only save money but can also prevent the occurrence of accidents.
- It can be prevented by processes like galvanising and electroplating with other metals.

Rancidity

- Oils and fats react with oxygen and get oxidised or turn rancid. This process is called rancidity.
- Rancidity can be prevented by keeping food in air tight containers or by using antioxidants.
- Antioxidants are used to prevent oxidation of food containing fats and oils.
- Storage of food in air tight containers also decelerates oxidation.

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