

Average Atomic Mass

- The atomic mass values on the periodic table are a weighted average of all the isotopes of a certain element. The weighting includes the % abundance and mass of the atom

Ex. Silver Ag-107 51.83 % abundance
 Ag-109 48.17 % abundance

For each isotope, divide the % abundance by 100 and then multiply by the mass number. Then add the result of each isotope together

$$\begin{aligned} \text{AAM} &= (107 \times 0.5183) + (109 \times 0.4817) \\ &= 107.97 \end{aligned}$$

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