

Note on Electrophilicity & Nucleophilicity

Nucleophile: Normally possesses a ~~lone~~ lone pair, or at least one π bond

Nucleophilic Attack



- nucleophiles donate an e^- pair to an electrophile to form a chemical bond
- nucleophiles are Lewis bases
- the more polarizable, more cloud, the more nucleophilic
 - $\rightarrow$ therefore, the tighter an atom holds its electrons, the less nucleophilic

Electrophile: is a reagent attracted to electrons

- either a positively charged species or a species having vacant orbitals
- electrophiles are Lewis acids
- they are positively charged, have an atom that carries a partial positive charge, or an atom that does not have a full octet