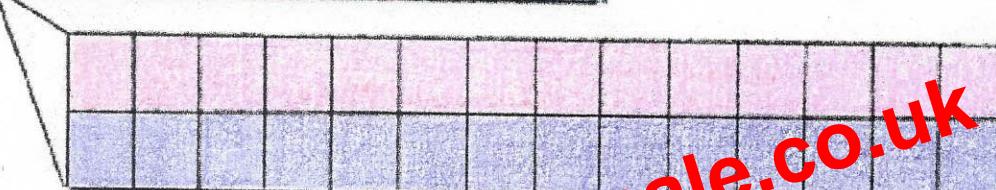


Metals are oxidized (lose e⁻)

Nonmetals are reduced (gain e⁻)

		Families and Oxidation Numbers										0	
		+1		+2		+3		-3		-2		-1	He
Not a metal ↗		H		alkali metals	noble gases								
Li	Be			alkaline earth	lanthanides								
Na	Mg			transition	actinides								
				halogens	Multiple Charges →								
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	+2		
Rb	Sr										+1 +2	+2 +4	Br Kr
Cs	Ba											+2 +4 Bi	I Xe Rn
Fr	Ra												



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Page 1 of 1

4

Elements in the same family have similar physical and chemical properties.

O.I.L. R.I.G.
oxidation reduction
e- e-

Alkali Metals:

- ~ most reactive metals
- ~ soft, silvery metals
- ~ catch fire if they come into contact with H₂O

Halogens:

- ~ most reactive nonmetals
- ~ combine with metals to form salts

Alkaline Earth Metals:

- ~ also reactive
- ~ found in compounds in earth's crust

Transition Metals:

- ~ typical metals
- ~ good conductors
- ~ ductile & malleable

Noble Gases:

- ~ all gases at room temp.
- ~ inert (unreactive)

Lanthanides:

- ~ rare earth metals

Actinides:

- ~ all are radioactive