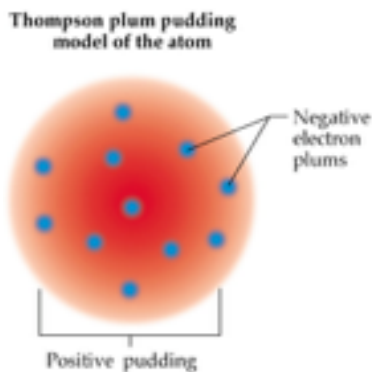
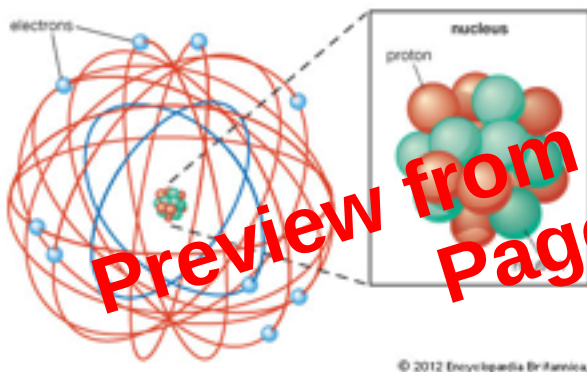


JJ Thomson: Pointed out that atoms are a uniform positive sphere which contain negatively charged particles called electrons and positively charged matter. He explained that electrons were like plums in a positively charged pudding thus it was called the 'plum pudding' model.



Rutherford: He reasoned that all of an atom's positively charged particles made up the nucleus. The negative charged particles in contrary were scattered outside the nucleus around the atoms edge.



Bohr: Bohr developed a model which showed that electrons didn't orbit randomly around the nucleus. Through his model he displayed that electrons moved with constant speed in fixed orbits around the nucleus. He proposed that they moved in specific layers at certain distances, and they could jump from a path on one level to another.

