

$$6.32 \text{ g KMnO}_4 = 6.32/158 \text{ moles KMnO}_4 = 0.04 \text{ moles KMnO}_4$$

$$\text{Concentration of the solution} = 0.04 \text{ mol l}^{-1} (= 0.04 \text{ M})$$

Question 52

Crystals of hydrated sodium carbonate ($\text{Na}_2\text{CO}_3 \cdot x\text{H}_2\text{O}$) of mass 3.15 g were dissolved in water and made up to 250 cm^3 in a volumetric flask. 25 cm^3 of this solution required 15 cm^3 of a