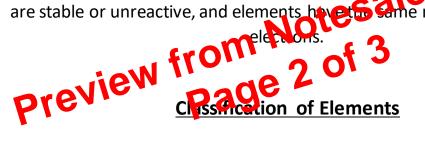
Periodic Table Organization

-The periodic table is arranged in order of increasing atomic number

-Elements are grouped according to similar properties

-Horizontal rows are called periods – within periods the atomic number increases, the electron cloud also increases within periods, and an example of this is that elements within period 3 have a larger electron cloud than elements in period 1.

-Vertical columns are called groups or families – within groups/family is elements have similar properties, elements form the same ions within groups, noble gases are stable or unreactive, and elements however same number of valence electors.



Metals – Make-up most the elements, they are located to the left of the metalloids, and Alkali metals are the most reactive metals

Non-Metals – Located to the right of the metalloids, and Halogens are the most reactive non-metals

Metalloids – Make-up the "staircase"/separation of metals from nonmetals