

Periodic Table Organization

- The periodic table is arranged in order of increasing atomic number
- Elements are grouped according to similar properties
- Horizontal rows are called periods – within periods the atomic number increases, the electron cloud also increases within periods, and an example of this is that elements within period 3 have a larger electron cloud than elements in period 1.
- Vertical columns are called groups or families – within groups/families elements have similar properties, elements form the same ions within groups, noble gases are stable or unreactive, and elements have the same number of valence electrons.

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Classification of Elements

Metals – Make-up most the elements, they are located to the left of the metalloids, and Alkali metals are the most reactive metals

Non-Metals – Located to the right of the metalloids, and Halogens are the most reactive non-metals

Metalloids – Make-up the “staircase”/separation of metals from non-metals