

Periodic table - Elements

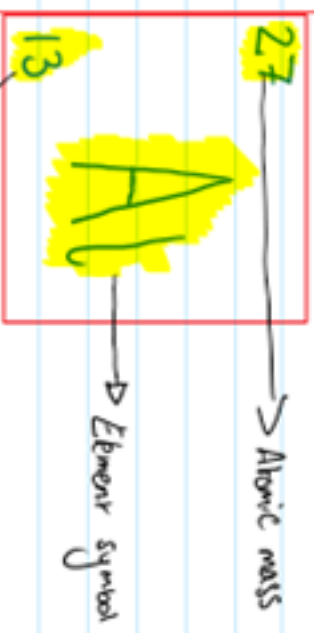
How to calculate:

N^o electrons:

Proton number = electron number

neutrons:

Atomic mass - atomic number = Neutron number



Atomic number

Isotopes: Element w/ same Atomic number but different Atomic mass.

Electronic Configuration of first 20 elements:

1	1	2	3	4	5	6	7	8	2
1	2,1	2,2	2,3	2,4	2,5	2,6	2,7	2,8	
2	2,1	2,2	2,3	2,4	2,5	2,6	2,7	2,8	
3	2,8,1	2,8,2	2,8,3	2,8,4	2,8,5	2,8,6	2,8,7	2,8,8	
4	2,8,8,1	2,8,8,2							

group
period

Formulae

Diatomic: F_2 I_2 Cl_2 H_2 N_2 O_2 Br_2
molecules: Fluorine Iodine Chlorine Hydrogen Nitrogen Oxygen Bromine

Compounds: $-ide$ eg. Iron Sulfide
w/ 2 elements Zinc Chloride

Compounds: $-ate$ eg. Zinc Sulfate
w/ oxygen Copper Sulfate

Ammonia: NH_3
Carbon monoxide: CO
Carbon dioxide: CO_2
Methane: CH_4
Sulfur dioxide: SO_2
Hydrochloric acid: HCl
Nitric acid: HNO_3
Sulfuric acid: H_2SO_4
Copper (I) oxide: CuO
Hydrogen oxide: H_2O
Magnesium hydroxide: $MgOH$
Calcium carbonate: $CaCO_3$
Copper carbonate: $CuCO_3$
Sodium chloride: $NaCl$
Sodium carbonate: Na_2CO_3
Copper (I) sulfate: $CuSO_4$

Charge	Proton	Neutron	Electron
Mass	+1	0	-1
	+1	+1	0.003