14.	ANS:	D PTS: 1 DIF: L2 REF: p. 222
	OBJ:	8.1.1 Identify the information a molecular formula provides.
	STA:	C.3.1 C.3.2 C.3.3 BLM: comprehension
15.	ANS:	A PTS: 1 DIF: L1 REF: p. 227
	OBJ:	8.2.1 Explain the result of electron sharing in covalent bonds.
	STA:	C.3.1 C.3.2 C.3.3 BLM: comprehension
16	ANS	A PTS 1 DIF I 1 RFF n 223
10.	OBI.	8.1.1 Identify the information a molecular formula provides
	STA.	$C_3 C_3 C_3 $
17	ANS.	C = PTS = 1 DEF L 2 REF n 226
17.	ORI-	8.2.1 Explain the result of electron sharing in covalent honds
	STA.	$C_{31} C_{32} C_{33}$ BIM: comprehension
18	ANS.	$\begin{array}{c} B \\ B \\ DTS \\ 1 \\ DIF \\ I \\ 2 \\ DF \\ I \\ $
10.	ORI-	8.2.1 Explain the result of electron charing in covalent honds
	STA-	$C_{3,1} = C_{3,2} = C_{3,3}$ BIM: comprehension
10	ANG.	$D = D = D = 1 $ DEv. L2 DEF. $\pi 220$
19.	ANS:	D PIS: I DIF: L2 REF: p. 230
	OBJ:	8.2.1 Explain the result of electron sharing in covalent bonds. O(2,1) + O(2,2) + O(2,2)
20	SIA:	C.5.1 C.5.2 C.5.5 BLW: comprehension
20.	ANS:	B PIS: I DIF: L2 REF: p. 227
	OBJ:	8.2.1 Explain the result of electron sharing in covalent bonds.
	STA:	C.3.1 C.3.2 C.3.3 BLM: application
21.	ANS:	D PTS: 1 DIF: L3 C REF: p. 226 p. 227
	OBJ:	8.2.1 Explain the result of electron sharp in ovarent bonds
	STA:	C.3.1 C.3.2 C.3.3
22.	ANS:	A PTS DIF: 1.2 REF: p. 226
	OBJ:	8.1.2 Descriptive representative units that define molecular compounds and ionic compounds.
	ST	A.C.N.32TC.3.3 DAOGBLM: comprehension
23.	AN S:	D PTS: DIF: L2 REF: p. 238 p. 244
	OBJ:	8.3.1 Describe the relationship between atomic and molecular orbitals.
	STA:	C.3.1 C.3.2 C.3.3 C.3.4 BLM: application
24.	ANS:	A PTS: 1 DIF: L2 REF: p. 230
	OBJ:	8.2.1 Explain the result of electron sharing in covalent bonds.
	STA:	C.3.1 C.3.2 C.3.3 BLM: application
25.	ANS:	B PTS: 1 DIF: L2 REF: p. 230
	OBJ:	8.2.1 Explain the result of electron sharing in covalent bonds.
	STA:	C.3.1 C.3.2 C.3.3 BLM: application
26.	ANS:	B PTS: 1 DIF: L1 REF: p. 224
	OBJ:	8.1.2 Describe the representative units that define molecular compounds and ionic compounds.
	STA:	C.3.1 C.3.2 C.3.3 BLM: knowledge
27.	ANS:	D PTS: 1 DIF: L2 REF: p. 232
	OBJ:	8.2.2 Describe how coordinate covalent bonds are different from other covalent bonds.
	STA:	C.3.1 C.3.2 C.3.3 BLM: comprehension
28.	ANS:	A PTS: 1 DIF: L2 REF: p. 233
-0.	OBJ:	8.2.4 Explain how the strength of a covalent bond is related to its bond dissociation energy.
	STA:	C.3.1 C.3.2 C.3.3 BLM: application
29	ANS	B PTS: 1 DIF: L_2 REF: n 236 ln 237
<i>_</i> /.	OBI.	8.2.4 Explain how the strength of a covalent bond is related to its bond dissociation energy
	STA.	$C_{31} C_{32} C_{33}$ BLM: application
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Network solids are substances in which all of the atoms are covalently bonded to each other. Melting these substances requires breaking covalent bonds throughout the solid. Two examples are diamond and silicon carbide.

PTS:1DIF:L2REF:p. 252OBJ:8.4.3 Explain why the properties of covalent compounds are so diverse.STA:C.3.1 | C.3.2 | C.3.3BLM: comprehension

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