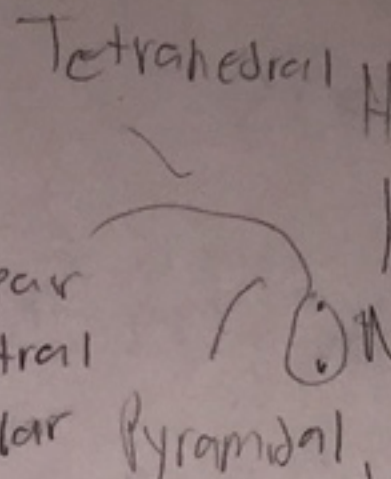
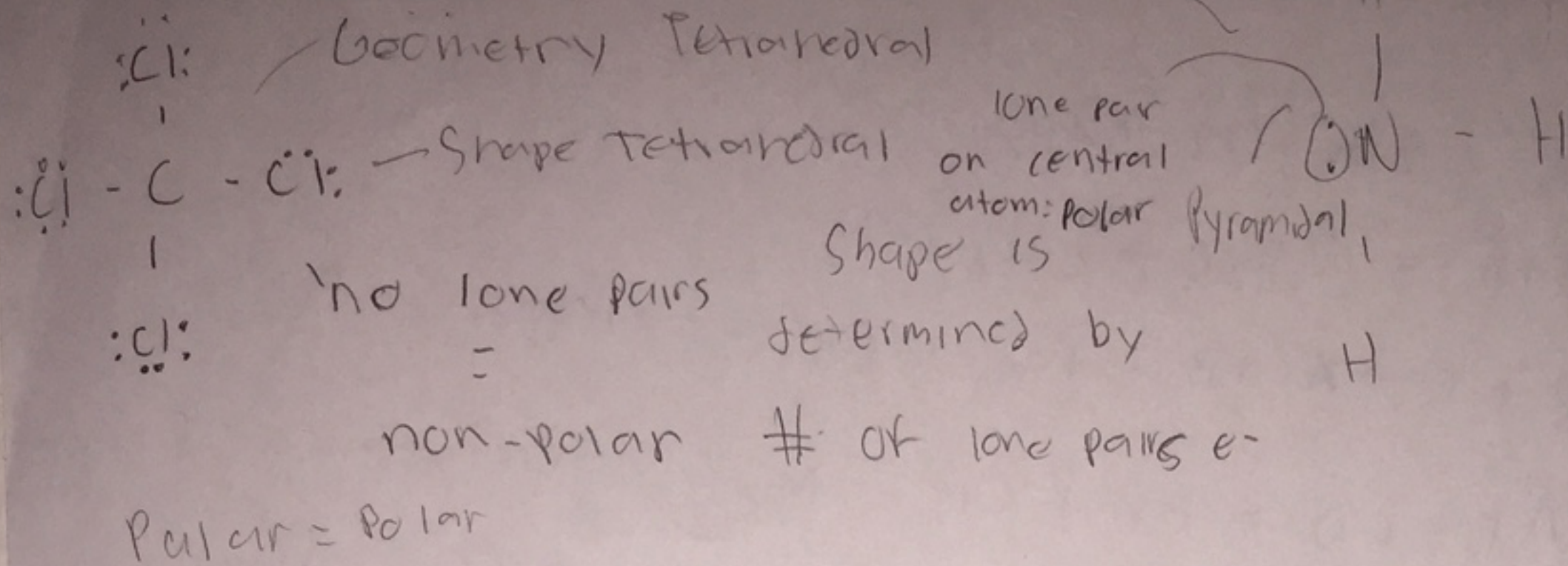
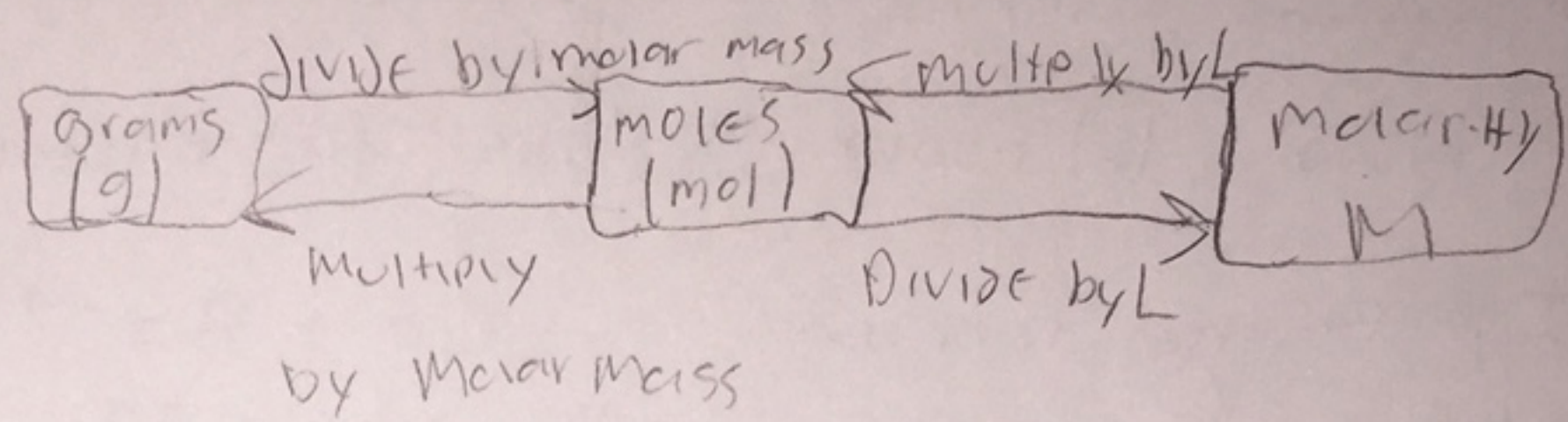


Molecules



Moles

1 mole = 6.02×10^{23}



molar mass - mass of 6.02×10^{23} particles

Molecular Formula

actual # of elements in a compound ex. C_2H_6

Empirical Formula

a formula that gives the simplest/lowest # ratio of atoms in a compound

Percent Composition

1. x molar mass of element by subscript of element in the compound
2. divide total molar mass of compound 40.9
3. multiply by 100

factors that affect Molarity/solubility - two substances being the same polarity

Solutions

mixture - two or more solutions
 Heterogeneous particles aren't evenly distributed
 Homogeneous particles are evenly distributed
 also know as a solution

\uparrow Pressure = \uparrow Solubility of gases in liquid

\uparrow temp = \uparrow Solubility of solids in liquids

$\text{MAcid} \rightleftharpoons \text{acid} = (\text{KAcid})(\text{LBase})$

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