

#catatan

NO Selasa
DATE 26 sept 2016

Periodic table

Chemical Bonding

Ionic Bonding
Covalent Bonding

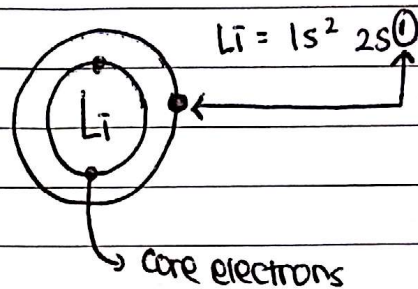
The alphabets of the universal

? why pure elements cannot be found freely in nature?

1	→ atomic number = Proton
H	→ symbol
Hydrogen	

Valence Electrons (elektron terluar)
-outermost-

• $\sum \text{proton} = \sum \text{electron} \Rightarrow \text{elements}$
↓
no charge



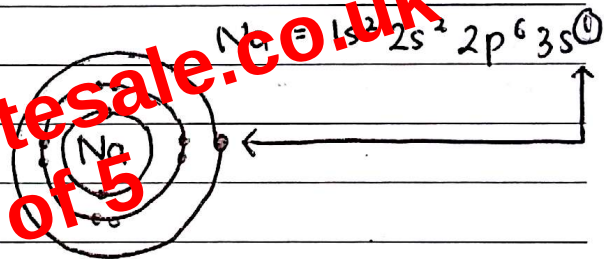
• $K^+ \rightarrow$ release one electron

① 7 rows

② 2 Groups → A & B (U.S.)

③ 18 Groups → periodic (1st, 2nd, 3rd, 4th, 5th, 6th, 7th, 8th, 9th, 10th, 11th, 12th, 13th, 14th, 15th, 16th, 17th, 18th)

ex: -13 - IIIA



→ semi metal
metal & nonmetal

Metal

Non-Metal

conduct of electricity

malleable → ditempa

Lewis dot symbol
(Berdasarkan golongan)

ex: 1A ... 8A → noble gas

H $\cdot$

Li $\cdot$

Na $\cdot$

He $\cdot\cdot$

Ne $\cdot\cdot$

etc...

-stabil-

Duplet structure

Octet structure

except He

(Noble Gas)

car: semi metal / semi conductor

- conduct electricity but
unmalleable

• $H_2O \rightarrow$ Bonding

Ionic Bonding

- release & receive

two or more atom

ex: NaCl

chemical bonding formed by
the electrostatic attraction
between $\oplus$ & $\ominus$ ions

release $\oplus$

receive $\ominus$