The two value represent the fact that the two electrons in an orbital have equal but opposite spin.

⇒ (i) PAUL'S EXLUSION PRINCIPLE → No two electrons with in the same atom can have the same set of all four Quantum number.

(ii) AUFBAU PRINCIPLE → Electrons are filled in the subshell in increasing order of the energies.

→ The order of energy is determined using the (n+l) rule.

(N+l) RULE :- (i) Subshell having smaller value of (n+l) is lower in energy

(ii) of the value of (n+l) is same the subshell having smaller value of 'n' will be lower in energy.

	1S	2S	2P	3S	3P	3d	4S	4p	4d	4f
(n+l)	1+0	2+0	2+1	3+0	3+1	3+2 ~~4+0~~	4+0	4+1	4+2	4+3
	=1	=2	=3	=3	=4	=5	=4	=5	=6	=7

Q.. Sc — Zn

$Sc_{21} = 1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^1$

$= [Ar]\ 4s^2 3d^1$

$Ti = [Ar]\ 4s^2 3d^2$

$V = [Ar]\ 4s^2 3d^3$

$*Cr = [Ar]\ 4s^1 3d^5$