

Atomic spectrum

like:

- 1.) For both the spectra, lines lie in the same position for a given element
- 2.) lines converge @ higher frequencies
- 3.) sets of lines represent transition from a particular energy level

Difference

Atomic spectrum
E.g. Emission spectrum



- 1.) Coloured lines on a black background
- 2.) lines converge at higher frequencies
- 3.) Energy gaps are specific to each element
- 4.) hence, compare X spectrum with known data to find corresponding element

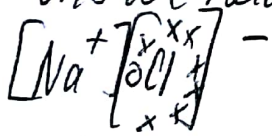
on coloured spectrum
on black background

Bonding
~ Covalent:

hydrogen is a metal

~ Ionic: Metals react with non metals

ions are held together via electrostatic bonds



~ Metallic bonding: $\begin{matrix} \oplus & \oplus & \oplus \\ \ominus & \ominus & \ominus \end{matrix}$ pool of delocalised electrons

Lattice of regularly spaced cations

~ Giant lattice (solid) [Ionic, covalent, metallic]

Structure

~ molecular (gas) [macromolecular, simple molecular]

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- 1.) melting
- 2.) boiling
- 3.) Electrical conductivity
- 4.) solubility in water
- 5.) solubility in non polar solvents