



Energy locked in chemical bonds

- When bonds are broken, energy is absorbed.
- When bonds are formed, energy is released.

■ Making or breaking weak bonds

(intermolecular forces) involve small energy changes.

Exothermic reactions (contd)

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- Reactants $\longrightarrow$ Products + heat energy

i.e. heat energy is transferred to the surroundings

Change in enthalpy

The reaction is exothermic if:

$$H_{\text{product}} < H_{\text{reactants}}$$

Heat energy is released to the surroundings,

ΔH is negative ($\Delta H - \text{ve}$)

Question (contd.)

- Draw a labeled energy profile diagram for the following reaction. Indicate on the diagram what effect adding a catalyst would have.





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