

Polyprotic Acids	An acid that can donate more than one proton or hydrogen atom per molecule to an aqueous solution. Examples: Sulfuric acid (H_2SO_4) is a polyprotic acid because it can donate two hydrogen atoms to an aqueous solution
Hydrolysis	a reaction involving the breaking of a bond in a molecule using water. The reaction mainly occurs between an ion and water molecules and often changes the pH of a solution. In chemistry, there are three main types of hydrolysis: salt hydrolysis, acid hydrolysis, and base hydrolysis
Buffer Solution	an aqueous solution consisting of a mixture of a weak acid and its conjugate base, or vice versa. ... Buffer solutions are used as a means of keeping pH at a nearly constant value in a wide variety of chemical applications
Hydrocarbons	a hydrocarbon is an organic compound consisting entirely of hydrogen and carbon, and thus are group 14 hydrides. Hydrocarbons, from which one hydrogen atom has been removed, are functional groups called hydrocarblys
Functional Group	a group of atoms responsible for the characteristic reactions of a compound
Alcohol Functional Group	an alcohol is any organic compound in which the hydroxyl functional group ($-\text{OH}$) is bound to a saturated carbon atom
Ethers Functional Group	An ether is an organic compound that contains two alkyl or aryl groups by an oxygen atom. The general formula for an ether is $\text{R}-\text{O}-\text{R}'$
Alkyl halides Functional group	The haloalkanes (also known as halogenoalkanes or alkyl halides) are a group of chemical compounds derived from alkanes containing one or more halogens. They are a subset of the general class of halocarbons, although the distinction is not often made
Amines Functional Group	The basic chemical structure is that of ammonia (NH_3) with the key atom being the central nitrogen atom. Try to remember that an amine is just like ammonia because ammonia is a simple molecule to recall. The basic ammonia structure is changed when the hydrogen atoms are replaced by alkyl groups to form amines
Alkanes	an acyclic saturated hydrocarbon. In other words, an alkane consists of hydrogen and carbon atoms arranged in a tree structure in which all the carbon-carbon bonds are single
Benzene	a colourless, volatile, flammable, toxic, slightly water-soluble, liquid, aromatic compound, C_6H_6 , obtained chiefly from coal tar: used in the manufacture of commercial and medicinal chemicals, dyes, and as a solvent for resins, fats, or the like
Esters	an organic compound made by replacing the hydrogen of an acid by an alkyl or other organic group. Many naturally occurring fats and essential oils are esters of fatty acids
Amides	an organic compound containing the group $-\text{C}(\text{O})\text{NH}_2$, derived from ammonia by replacement of a hydrogen atom by an acyl group
Aldehydes	an organic compound containing a functional group with the structure $-\text{CHO}$, consisting of a carbonyl centre (a carbon double-bonded to oxygen) with the carbon atom also bonded to hydrogen and to an R group, which is any generic alkyl or side chain
Ketones	an organic compound with the structure $\text{RC}(=\text{O})\text{R}'$, where R and R' can be a variety of carbon-containing substituents. Ketones and aldehydes are simple compounds that contain a carbonyl group (a carbon-oxygen double bond)
Carboxylic Acids	any of a class of organic compounds in which a carbon (C) atom is bonded to an oxygen (O) atom by a double bond and to a hydroxyl group ($-\text{OH}$) by a