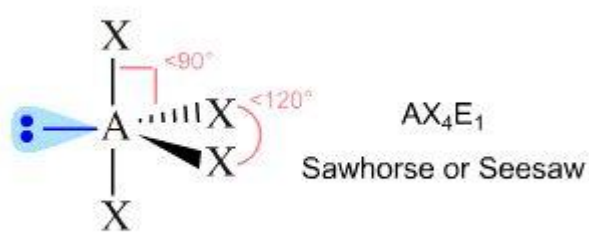
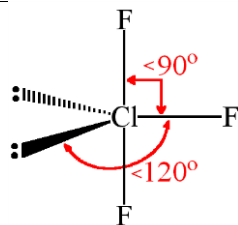
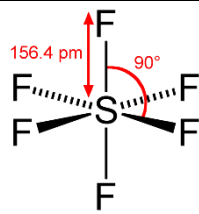
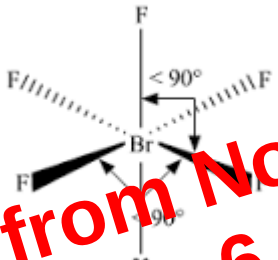
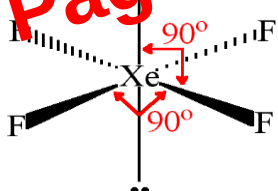


4BP 1LP		seesaw
3BP 2LP		T-shape
6BP		Octahedral
5BP 1LP		Square pyramidal
4BP 2LP		Square

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E.g. H₂O

H: 2* (1 electron)

2 bonds, seen from H₂O, no double bonds as Hydrogen isn't -vely charged

O: 6 electrons

Total= 8 electrons

2BP = 4 electrons (need a total of 8 electrons) so add 2LP = 8 electrons

2BP 2LP = bent

