

materials to form isolated atoms in the gas phase, followed by recombination of the atoms to form the products of the reaction. This is an easier way to watch signs.

7.14 Using Enthalpies of Atom Combination to Probe Chemical Reactions

Isomers = same formula, different structures

7.15 Bond Length and the Enthalpy of Atom Combination

Strength of bonds is related to bond length

As bond length increases, bond strength decreases

7.16 Hess's Law

Hess's Law states that ΔH_{rxn} is the same **regardless of whether a reaction occurs in one step or in several steps**, and **regardless of the path** by which the **reactants are converted into the products** of the reaction.

7.17 Enthalpies of Formation

Enthalpy of Formation = the enthalpy of reaction for each of the reactions = ΔH_f°

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