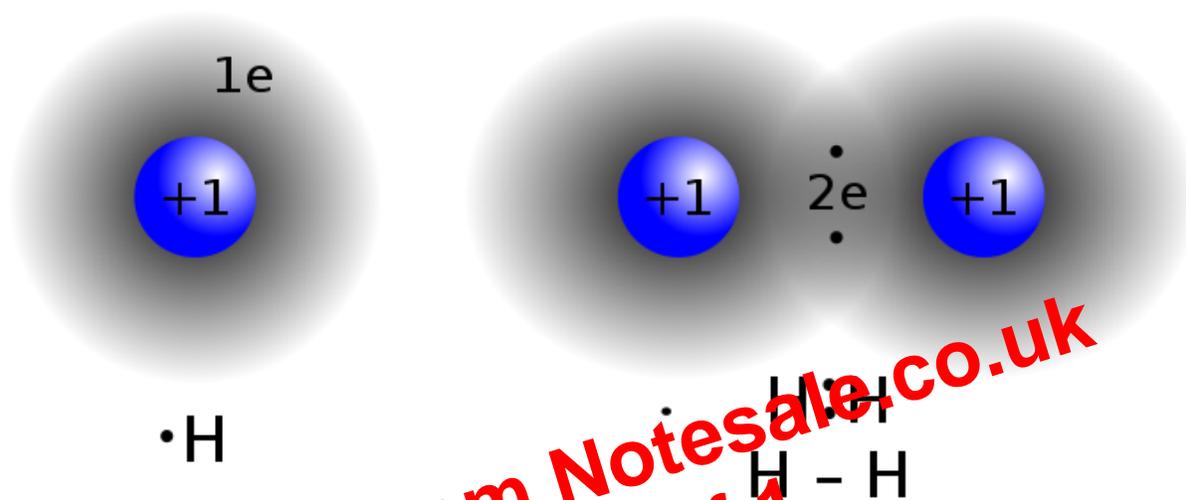


Compound formed by sharing electrons are known as **COVALENT COMPOUND**.

Example:

Formation of hydrogen molecules ( $H_2$ ) hydrogen atom contains one electrons in its outer most shell. In order to form hydrogen molecules both atoms share one electron each so as to form compound properties of two electrons.



### PROPERTIES OF COVALENT COMPOUND

- Consist of molecules (not ions)
- Are usually gases or liquid with low boiling points and melting point.
- No electrolytes (do not conduct electricity)

### DIFFERENT BETWEEN ELECTROVALENT AND COVALENT BONDING

ELECTROVALENT	COVALENT
Are crystalline solid	Most of them are liquid
Have high melting point and boiling point	Have low melting point and boiling point
Are insoluble inorganic solvent	Are soluble inorganic solvent
Consist of ions; eg $Na^+$ , $Cl^-$	Consist of molecules; eg $H_2$