

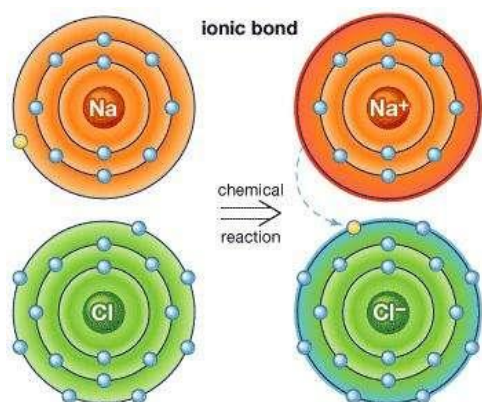
The sodium (Na^+) and chlorine ion (Cl^-) attack each other by an electrostatic force and the bond between them is ionic or electrovalent.

Before combination

$\text{Na}^+ = 2:8:1$

$\text{Cl}^- = 2:8:8:7$

Or



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- Other example of electrovalent compound is MgS ,

The electron gained or lost by one atom of an element during chemical combination so as to obtain a stable structure is called **ELECTROVALENT**.

PROPERTIES OF ELETROVALENT OR IONIC COMPOUND

- Ionic compound do not exist in molecules
- They are electrolytes conduct electricity when are molten state not solid
- They has high melting and boiling points
- Most of them are soluble in water but but insoluble in organic compound

I. COVALENT BONDING

Is a bond formed due to equally sharing of electrons from each atom to gain a stable structure