

	1. NaOH # reduced and oxidized * disproportionation	if heptane is added : upper layer turn purple (iodine in non-aqueous solvent) (NaOH) cold dilute $\text{NaCl} + \text{OCl}^-$ (sodium hypochlorite) Hot concentrated $\text{NaCl} + \text{ClO}_3^-$ (sodium chlorate)
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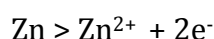
Reducing agent

Sulphite ions (SO_2 dissolve in H_2O - --- Sulphurous acid H_2SO_3 ---- SO_3^{2-}	$\text{SO}_3^{2-} \rightarrow \text{SO}_4^{2-}$	React with : 1. bromine 2. acidified KMnO_4 3. acidified $\text{K}_2\text{Cr}_2\text{O}_7$
Sulphur dioxide	$\text{SO}_2 \rightarrow \text{SO}_4^{2-}$	

Redox in chemical cell

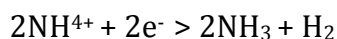
Zinc carbon cell – 1.5V

Cathode (negative) – zinc cup



electrolyte – ammonium chloride (NH_4Cl)

ammonium ions



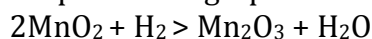
Hydrogen gas accumulate on the surface of graphite

- slows down + decrease the current produced

Anode (positive) – graphite

– surrounded by manganese (4)oxide (MnO_2)

– powdered graphite- increase electrical conductivity



reaction at graphite:

overall reaction :