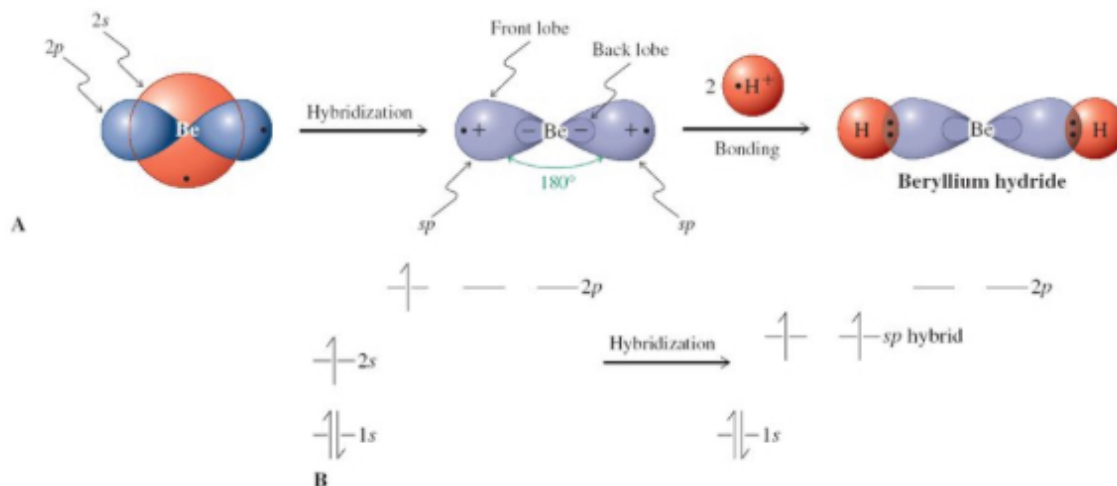




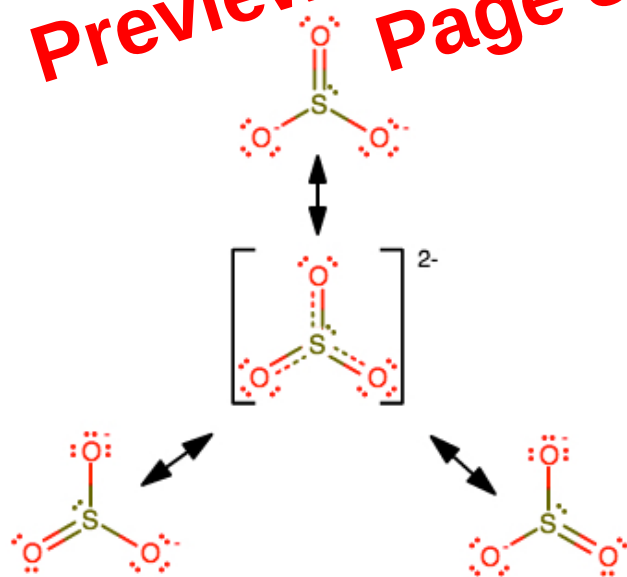
Note: The 2p, 2s labeling before hybridization; the front and back lobes after hybridized; the bond angle of 180 degrees; the 2 before the H 2s orbital <that is in the process of bonding>; the two dots that are in the bonded molecule

→ Sp^2 hybrids – Trigonal Structures:

The 3 hybrid orbitals formed by the mixture of one s and two p orbitals

The small amount of energy needed to promote an electron from the 2s to one of the 2p levels

SO₂



→ Sp^3 hybrids:

The 2s orbital is mixed with all three of the 2p orbitals, creating four hybridized sp^3 orbitals