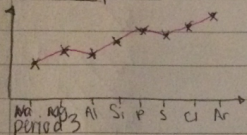


Trends in ionisation energy

Preview from Page

The more electrons you remove the more energy is required

Across a period



I.E. generally increases across a period.
because

- shielding is constant
- nuclear charge increases
- atomic radius decreases

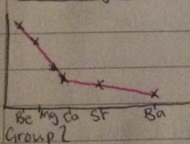
There are anomalies because

(Mg $\rightarrow$ Al) the outer electron is in a different sublevel

(P $\rightarrow$ S) in Phosphorus the electrons are all in different suborbitals

3p- In Sulfur one 3p sub orbital has 2 electrons which experience mutual repulsion -

Down a group



I.E. generally decreases down a group
because

- shielding increases
- nuclear charge increases

Copper and Cobalt are exceptions in electron structure