

Boiling Point

- > vapor pressure of liquid = atmospheric pressure
- > normal boiling point : temp. at 760 mm Hg
- > $\downarrow$ pressure, $\downarrow$ temp. of boiling point

Critical Temp & Critical Pressure

- > critical point - where the liquid reached critical temperature (high enough temp.) & critical pressure (high enough pressure)
- supercritical fluid

Surface Tension : energy required to break through the surface

- water drops to be spheres

Capillary action : rising of water due to attraction of two substances

- meniscus

- adhesive forces - attraction between liquid & non-liquid (solid)
- cohesive forces - water molecules themselves

Viscosity : resistance of liquids to flow

- Honey has long chains of molecules $\Rightarrow$ greater intermolecular forces