

Examples of reducing agents

- ① Metals
- ② Carbon
- ③ Hydrogen (H_2)
- ④ Iron(II) (Fe^{2+})
- ⑤ Alcohol
- ⑥ Halides [I^- , Br^- , I^- , At^-]
- ⑦ Sulphur dioxide (SO_2).

Oxidation numbers:

Grp I elements in a compound +1

Grp II elements in a compound +2

Grp III elements in a compound +3

Grp IV elements in a compound +4

Grp V elements in a compound +5

Grp VI elements in a compound +6

Grp VII elements in a compound -1

Grp VIII elements in a compound 0

⊖ Oxidation number of oxygen is -2, but in peroxide oxidation number is -1

⊖ Oxidation number of hydrogen is +1 but in metal hydride oxidation number is -1.

⊖ Oxidation number of any elements, molecules, compounds is zero 0.

⊖ Oxidation number of any ion is the charge on the ions.

$NO_3^- \rightarrow -1$	$MnO_4^- \rightarrow -1$	$CO_3^{2-} \rightarrow -2$
$OH^- \rightarrow -1$	$Cr_2O_7^{2-} \rightarrow -2$	$NH_4^+ \rightarrow +1$
$PO_4^{3-} \rightarrow -3$	$SO_4^{2-} \rightarrow -2$	$SO_3^{2-} \rightarrow -2$