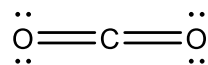
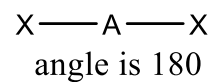


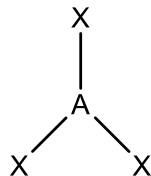
1. determine # of valence e-
2. draft put an initial structure with a center atom
3. apply the geometries to compounds with lone pairs
4. Group 7 will make one bond if not expanded octet
5. Boron will have a reduced octet (6 e- max)

Preview from Notesale.co.uk
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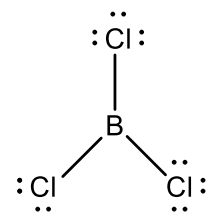
Linear

$$\text{CO}_2 = 4\text{e}^- + 12\text{e}^- = 16\text{e}^-$$



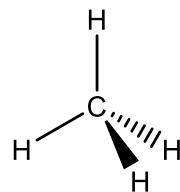
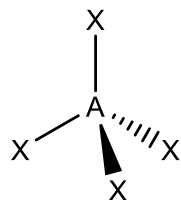
angle 120

$$\text{BCl}_3 = 3\text{e}^- + 21\text{e}^- = 24\text{e}^-$$



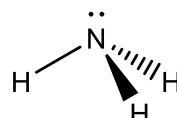
exception
6 e-

trigona planar



109.5

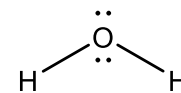
$$\text{NH}_3 = 8\text{e}^-$$



<109.5

trigonal pyramidal

$$\text{H}_2\text{O} = 8\text{e}^-$$



<<109.5

bent

tetrahedral