

↳ This results in a lone-pair repelling more strongly than a bonding pair.

Increasing repulsion →

bonded-bonded < bonded-lone < lone-lone

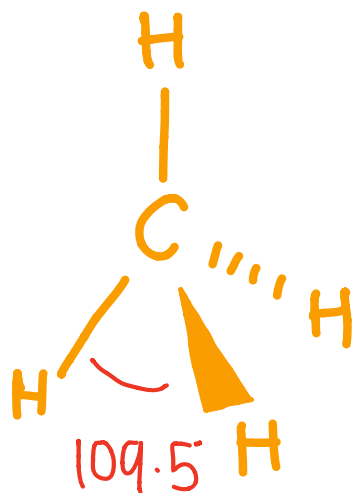
Four electron pairs:

ammonia and water → have a mixture of bonded-lone pairs → lone pairs repel more strongly → ∴ lone pairs repel bonded pairs slightly closer together → decreasing bond angle → bond angle decreased by 25° per lone pair.

In molecules containing multiple bonds, each multiple bond is treated as a bonding region.

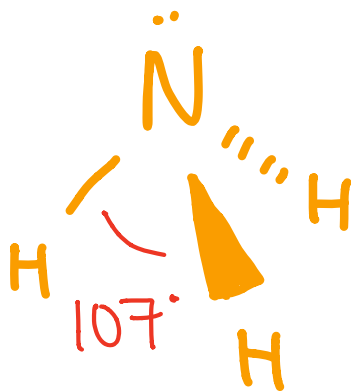
~~Tetrahedral~~

4 Bonded pairs
0 lone pairs



~~Pyramidal~~

3 bonded pairs
1 lone pair



~~non-linear~~

2 bonded pairs
2 lone pairs

