Total relative uncertainty = 0.0021186441 + 0.0047393365 + 0.0010288066 + 0.0120481928 == 0.0199349800

DATA PROCESSING AND PRESENTATION:

Equation of the reaction:

$$Mg + 2 HCl \longrightarrow MgCl_2 + H_2$$

pV = nRT

$$R = \frac{pV}{nT}$$

$$n(H_2) = n(Mg) = \frac{m(Mg)}{M(Mg)}$$

$$R = \frac{pVM(Mg)}{Tm(Mg)}$$

$$R = \frac{97.2*0.0498*24.305}{294.3*0.0472} = 8.475283024 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$$

Absolute uncertainty = $0.0199349800 * 8.475283024 = 0.1689545976 \text{ J·K}^{-1} \cdot \text{mol}^{-1}$ R = $8.5 \pm 0.2 \text{ J·K}^{-1} \cdot \text{mol}^{-1}$ Taking into consideration the vapuul passure of water:

the burette was a coequence of H₂, because there was probably some water vipour. Water vapour pressure ac 20 c (similar temperature as in the experiment) is around 2.3 kPa. Taking this into consideration, the pressure of H₂ was a little bit lower than the room pressure, which leads to smaller value of R constant.

$$R = \frac{(p-2.3kPa)VM(Mg)}{Tm(Mg)}$$

$$R = \frac{(97.2 - 2.3) * 0.0498 * 24.305}{294.3 * 0.0472} = 8.269112869 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$$

Absolute uncertainty = $0.0199349800 * 8.269112869 = 0.1648445997 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$

$$R = 8.3 \pm 0.2 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$$

When taking into consideration the vapour pressure of water the final result was more accurate.

CONCLUSION AND EVALUATION:

The final result was quite accurate and precise, specially after taking into consideration the vapour pressure of water. Experimentally obtained gas constant was $8.3 \pm 0.2 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1}$, while the actual value of R is 8.31 J·K⁻¹·mol⁻¹. The actual value is within the range of the experimental values including uncertainties.