n=1 → l=0 →s-subshell
n=2 → l=0,1 →s,p-subshell
n=3 → l=0,1,2 →s,p,d-subshell
n=4 → l=0,1,2,3 →s,p,d,f-subshell

3: Magnetic Quantum number "m"

m=(2l+1) or **m=+l,0,-l**

Gives information about the orbitals.

- No of electrons in one orbital=2
- In magnetic field orbitals degenerate,(dxy,dyz,dzx,dx^2-dy^2,dz^2)

l=0=s , m=0
l=1=p , m=3 or **m=+1,0,-1**
l=2=d , m=5 or **m=+2,+1,0,-1,-2**
l=3=sf , m=7 or **m=+3,+2,+1,0,-1,-2,-3**

4: Spin Quantum number "s"

s=+1/2↓(clockwise) **s=-1/2↑(anticlockwise)**

- Due to the opposite magnetic field, the electrons don't repel each other, rather attract each other.