constant or also called as the ebullioscopic constant. Molal elevation constant K_b have the unit that is equivalent to K kg mol^{-1}.
Suppose that m_1 and m_2 are the masses of solvent and solute while the molar masses of solute is equal to M_2 and the solvent is equal to M_1, then the molality is equivalent to;

$$m = \frac{\frac{m_2}{M_2}}{\frac{m_1}{1000}} = \frac{1000 \times m_2}{M_2 \times m_1}$$

By put the value of molality in the boiling of elevation equation then;

$$\Delta T_b = \frac{k_b \times 1000 \times m_2}{M_2 \times m_1}$$

$M_2 = K_b \times 1000 \times m_2 / \Delta T_b \times m_1$

So, by knowing the others quantities we can calculate the molar mass of solute of the elevation of boiling point equation.

- **Example:**

Pure liquid have the boiling point that is equivalent to 353.23 K. When dissolve the 2.70 g of the non-volatile solute in to the liquid (90g), then solution' boiling point increased to 354.11 K. Calculate the molar mass of the solute? By using the value of K_b of liquid that is equal to 2.53 K kg mol^{-1}.

Since that;

$T_b^o = 353.23$ K
$T_b = 354.11$ K
$m_2 = 2.70$ g
$K_b = 2.53$ K kg mol^{-1}
and $m_1 = 90$ g
Elevation of boiling point is equivalent to;

$\Delta T_b = T_b - T_b^0$
$= 354.11K - 353.23K = 0.88$