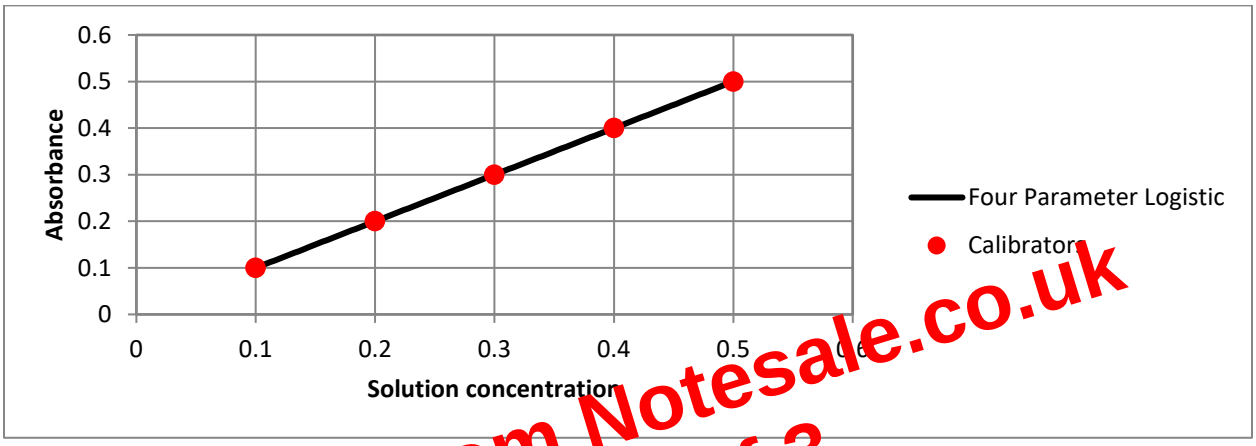
SAMPLE #1: Concentration of $Cu(NO_3)_2$ in measured sample (from calibration curve), mole/L	Concentration= (0.350/1.2365) Concentration= 0.283 mol/L
SAMPLE #2: Concentration of $Cu(NO_3)_2$ in measured sample (from calibration curve), mol/L	Concentration= (0.567/1.2365) Concentration= 0.459 mol/L



Discussion / Conclusion

Utilizing Beer's, not really set in stone the worth of ε · ℓ by tracking down the worth of A with arrangements of known fixation C This worth was then used to compute C of our obscure example from its deliberate worth of A. An alignment bend was made to get a precise worth of ε · ℓ. This is a straight-line bend that goes through the beginning with a steady equivalent to 0 and an incline of ε · ℓ, which is 1.2365. The condition of the line is y= 1.2365x, where y is the absorbance worth and x is the arrangement focus. By revamping the condition, the molar grouping of the obscure metal particle can be determined by isolating the absorbance esteem by 1.2365. The primary example molar focus was determined by partitioning 0.350 by 1.2365. This brought about a molar convergence of 0.283 mol/L for test #1. The second example molar fixation was determined by partitioning 0.567 by 1.2365. This brought about a molar centralization of 0.459 mol/L for test #2. This equivalent research facility technique is a decent method for deciding the convergence of NaCl in water since it is dull and doesn't ingest noticeable light.