

Class - XI

Some Basic Concepts of Chemistry

(i) Eq. wt, $E = \frac{\text{wt given}}{\text{No. of equivalents}}$

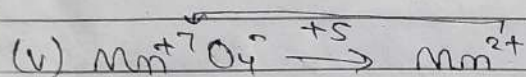
(ii) $E = \frac{A}{V}$ [$E = \text{eq. wt}$, $A = \text{at. mass}$, $V = \text{valency}$]

(iii) Dulong Petit's Law — $\text{At. wt} \times \text{sp. heat} = 6.4$
(approx)

(iv) Eq. wt of anions = $\frac{\text{Mol wt}}{\text{Charge}}$

Eq. wt of salt = $\frac{\text{Mol wt}}{\text{No. of cations} \times \text{Charge}}$

(No. of cations is 1st preference for compounds)



Eq. wt = $\frac{\text{Mol wt}}{\text{Change in oxidation no.} \times \frac{1}{5}}$
(m-factor)

(vi) Normality, $N = \frac{\text{No. of equiv of solute}}{\text{Vol. of sol (ml)}} \times 1000$

(vii) Eq. wt of acid = $\frac{\text{Mol wt}}{\text{Basicity of acid}}$

(viii) $\text{ml} \times (N) = \text{milli equiv}$; $\text{ml} \times (M) = \text{milli mol}$;

eq. tension = Vapour Pressure (VP);

$1(N) = \frac{1}{2}(M)$ — for di-basic acid

(ix) Final strength = $\frac{\text{wt given} \times 1000}{\text{Eq. wt} \times \text{Vol of sol}}$
No. of equiv