Ans. In the formation of anion , one or more electrons is added in outermost shell. In doing so, the no. of shells remain same but the no. of cation per electron decreases. As a result, repulsion of atomic size increases as compound to their atom.

Ques. **Why the size of atom of inert gases is large in their particular period?**

Ans. In the case of inert gases , the inner electronic shells are filled and due to repulsion they expands.As a result, their size increases and also their atomic radius is calculated in the turns of Vander Wall Radius and Vander Wall Radius is always greater than the covalent radius.

CONFIRMATION OF GROUP AND PERIOD—

PERIOD—The no. of shells in doing the electron configuration of any atom tellsabout the no. of period to which that atom belong.

EXAMPLE;

He= 2-K = 1st period

Li= 3[2,1]-K,L= 2nd –period

Na= 11[2,8,1] –K,L,M = 3rd –period

K= 19[2,8,8,1]-K,L,M,N = 4th –period

GROUP—The no. of electrons present in the valence shell gives the information about the group no.

NOTE..

1. If 1 and 2 electrons are present in valence shell then the atom belongs to either 1st group or 2nd group.