

Father Of Modern Chemistry - Lavoiser

Matter - Anything that occupies space, possesses mass.

Solids - definite shape, definite volume, cannot be compressed, hard and rigid, eg. **Metals, wood, bricks, copper, etc**

Liquid - definite volume, no definite shape, eg. **Water, milk, oil, alcohol, etc**

Gaseous - no shape, no volume, eg. **Air, nitrogen, oxygen, hydrogen, etc**

Pure substances - A single Substances which cannot be separated by any physical process. Classified as **Elements and Compounds**.

Elements - simplest form of pure substance which can neither be broken nor built from simpler substances by physical methods. Three types of Elements: **Metals, Non metals, Metalloids**

Metals - solids. (Exception mercury - it is liquid at room temperature) have Lustre, high Melting point, high Boiling point, good conductor of heat and electricity. Conductivity of metal decreases with increase in temperature due to vibration of positive ions at lattice points. Eg. **Iron, copper, zinc, silver, gold, aluminum**

Non metals - its found in all states of matter. Do not possess Lustre (exception iodine), poor conductor of heat and electricity (exception graphite). Not malleable and ductile. Eg. **Hydrogen, oxygen, nitrogen, Sulphur, phosphorous, etc**

Metalloids - common properties of metals and non metals eg. **Arsenic, antimony, bismuth, etc**

Compounds - composed of two or more elements in fixed proportion by mass. Properties are different from elements. Eg. **Water, sugar, salt, chloroform, alcohol, ether, etc**. Classified into **Organic and inorganic compounds**