7. Silver has two naturally occurring isotopes ^{107}Ag (106.90509 amu) and ^{109}Ag (108.90475 amu). The average atomic mass of silver is 107.8682 amu. The fractional abundance of ^{107}Ag is:

A) 0.3817
B) 0.7578
C) 0.2422
D) 0.5184
E) 0.9047

Part II.

8. Complete the following table (all empty cells):

NAME	FORMULA
Aluminum sulfate	$Al_2(SO_4)_3$
nitrogen triiodide	NI_3
Magnesium carbonate	$MgCO_3$
manganese (IV) oxide	MnO_2

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9. Complete the following table (all empty cells in both rows):

Symbol	# of protons	# of neutrons	# of electrons	atomic number	mass number
^{65}Cu	29	36	28	29	65
^{19}F	9	10	9	9	19

10. Maleic acid, which is used to manufacture artificial resins, has the empirical formula CHO. Its molar mass is 116.1 g/mol. What is its molecular formula?

$CHO = 12.011 + 1.008 + 16.00 = 29.019$ g/mol

$\frac{116.1}{29.019} = 4$ ∴ molecular formula =

$4 \times (CHO) =$ $C_4H_4O_4$