Initial buret reading (**mL**)	**7.60** Edit	**5.85** Edit
Final buret reading (**mL**)	**23.55** Edit	**24.95** Edit

Lab Calculations

Standardization of NaOH

Volume of base used (**mL**)	Try 1: ✔ 15.3 3 3 Info	Try 1: ✔ 15.55 3 3 Info
Molarity of the NaOH solution (**M**)	Try 1: ✘ 0.02089 Try 2: ✔ 0.2089 6 6 Info	Try 1: ✔ 0.2055 6 6 Info
Average Molarity of NaOH (**M**)		Try 1: ✔ 0.2072 3 Info

How is the standardization of NaOH impacted by the balance you used? Explain it. Refer to which balance you used to weigh out the KHP? 10 10

I noticed that the balance I used was very sensitive to even the slightest bump on the table. It is possible that my numbers for the amount if KHP I had were slightly skewed if I made a slight bump on the table.

Edit Answer

[Excellent! You mentioned the decimal places and the sensitivity/precision of the balance. You also related it to the precision of the buret.]

Instructor Assigned Grade Based On The Experimental Molarity Of NaOH 10 10 View Comment

Titration of the Diluted HCl

	Trial 1	Trial 2
Volume (**mL**) of acid used	9.25	8.75
Initial buret reading (**mL**)	7.60	5.85
Final buret reading (**mL**)	23.55	24.95