NaOH Concentration (**M**)	**0.2105** Edit

Try 1: ✘ 204
Try 2: ✔ 0.0071

Volume (**in L**) of a 6M HCl stock solution needed to make 100 mL of a 0.43 M HCl solution

10 | 10
Info

Does it make a difference if you use a volumetric flask vs a beaker vs a graduated cylinder to make your solution? Explain.

20 | 20

The graduated cylinder is used for measuring out liquids where as the volumetric flask is better suited to make the needed solution

Edit Answer
[Excellent! You discussed the accuracy and/or precision of each glassware AND what each was best suited for]

What amount, in mL, of acid did you **actually** put in the volumetric flask to make your 0.43 M solution	**5** Edit

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What did you use to measure the volume of acid and why?

10 | 10

I used the volumetric pipette as it had 5 ml marked on it which is what I needed to get.

Edit Answer
[Excellent! You chose a piece of glassware that provided the needed precision and explained]

LAB DATA

Determining Concentration of Unknown Solution

	Trial 1	Trial 2	Trial 3
Volume of ~0.43 M acid used aka take a small volume OUT of the volumetric flask.	**25** Edit	**10** Edit	**10** Edit
Initial buret reading (**mL**)	**24.85** Edit	**6.25** Edit	**24.55** Edit
Final Buret reading (**mL**)	**34.35** Edit	**9.25** Edit	**48.45** Edit

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Lab Calculations

	Trial 1	Trial 2	Trial 3
	Try 1: ✔ 9.5	Try 1: ✔ 3	Try 1: ✔ 23.9
Volume added (**mL**)	3 | 3 Info	3 | 3 Info	3 | 3 Info
