

Atomic structure

recap

Particle	rel. mass	rel. charge	actual mass / g	actual charge / C
proton	1	+1	1.67×10^{-24}	$+1.602 \times 10^{-19}$
neutron	1	0	1.67×10^{-24}	0
electron	$\frac{1}{1840}$	-1	9.10×10^{-28}	-1.602×10^{-19}

→ Relative scales are used because they are simple

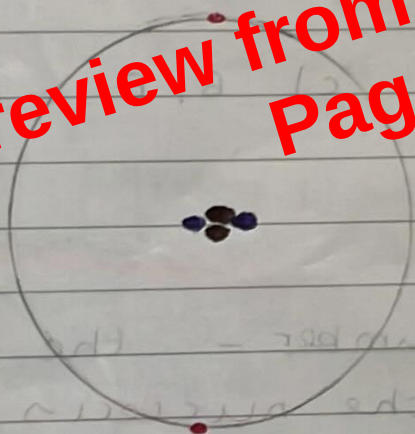
→ $\frac{1}{12}$ of a carbon-12 atom

Size: Oxygen = 7.4×10^{-11} m

3.311×10^{25} atoms in 50 ml of water

(drinks can) mostly empty space

Preview from Notesale.co.uk
Page 1 of 1



• Electron

Nucleus →

• Protons

• neutrons

} nucleus

Nucleons are subatomic particles with a relative mass of 1, which are found in the nucleus